

# Secrets of a university tutor: tetrahedral intermediates.

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## Citation

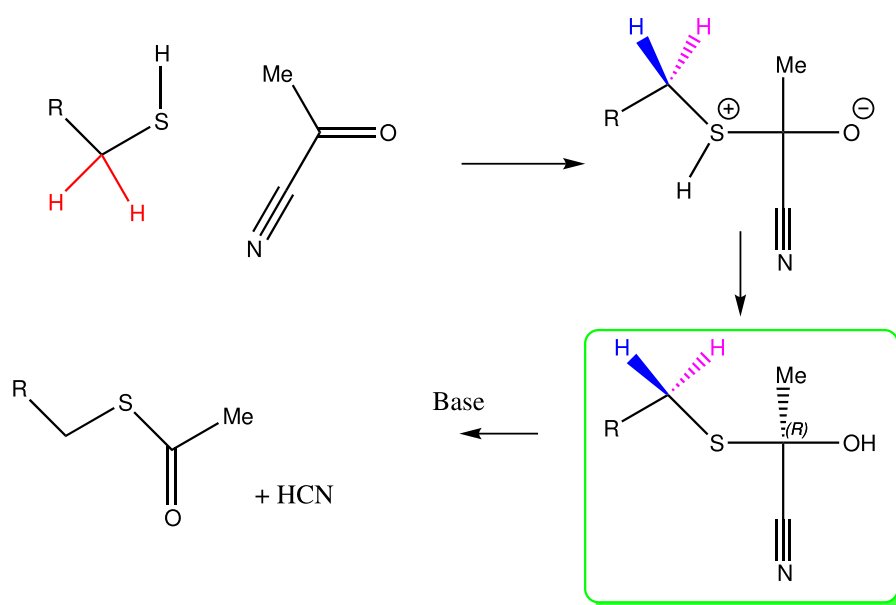
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## Keywords

Interesting Chemistry, Reaction Mechanism, Diastereotopic, Spectroscopy, Tetrahedral Intermediate

## Abstract

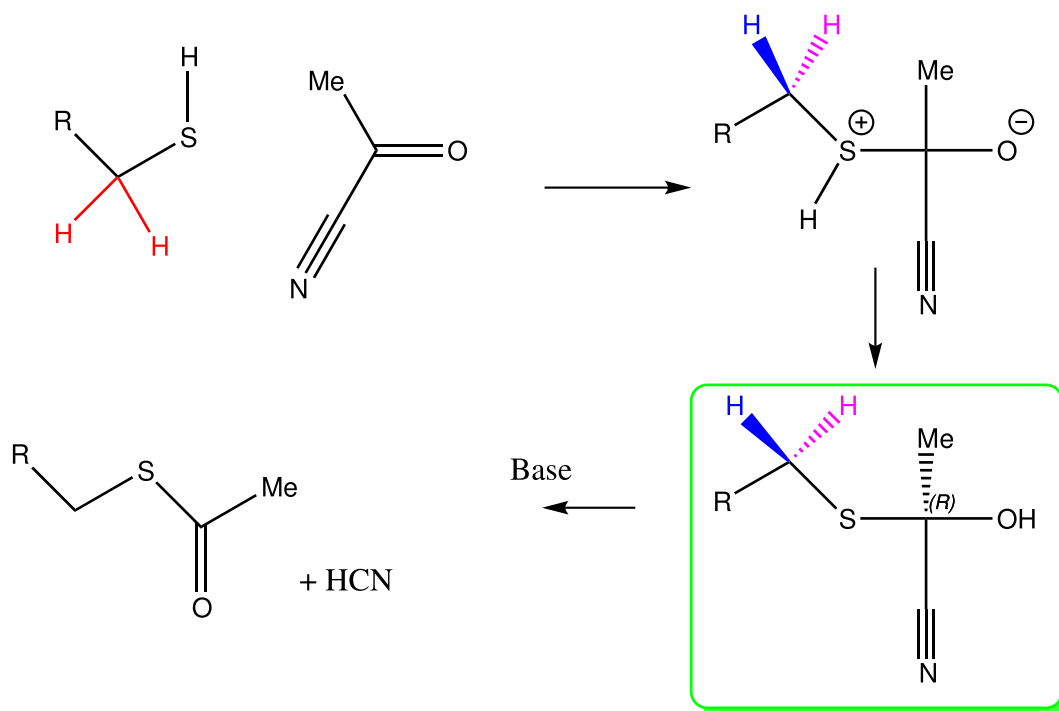
The tetrahedral intermediate is one of those iconic species on which the foundation of reaction mechanism in organic chemistry is built. It refers to a (normally undetected and hence merely inferred) species formed initially when a nucleophilic reagent attacks a carbonyl compound. Its importance to understanding the activity of enzymes cannot be overstated.



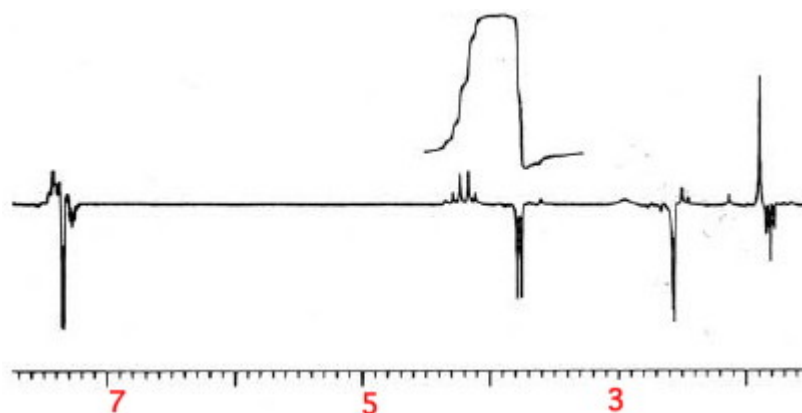
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The **tetrahedral intermediate** is one of those iconic species on which the foundation of reaction mechanism in organic chemistry is built. It refers to a (normally undetected and hence merely inferred) species formed initially when a nucleophilic reagent attacks a carbonyl compound. Its importance to understanding the activity of enzymes cannot be overstated. An example of this genre is shown below, in which a thiol reacts with an acyl cyanide to form the species ringed in green.



The lifetime of such a species is normally assumed to be too short for it to be detectable[cite]10.1021/jo00389a050[/cite], but suitable adjustment of the substituents enables its lifetime can be extended to many hours. Thus the reaction between  $\text{PhCH}_2\text{SH}$  (benzyl thiol) and acetyl cyanide occurs slowly and the resulting intermediate is stable enough to be detected by  $^1\text{H}$  NMR spectroscopy[cite]10.1021/ed064p725[/cite]. The spectrum in  $\text{CDCl}_3$  is shown below:



Spectrum of a tetrahedral intermediate.

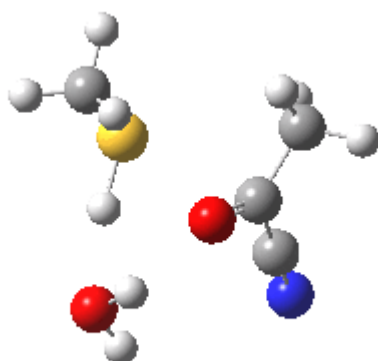
What is shown is actually a difference spectrum, in which the spectrum measured immediately after mixing the two reagents is subtracted from that obtained after about 12 hours. The

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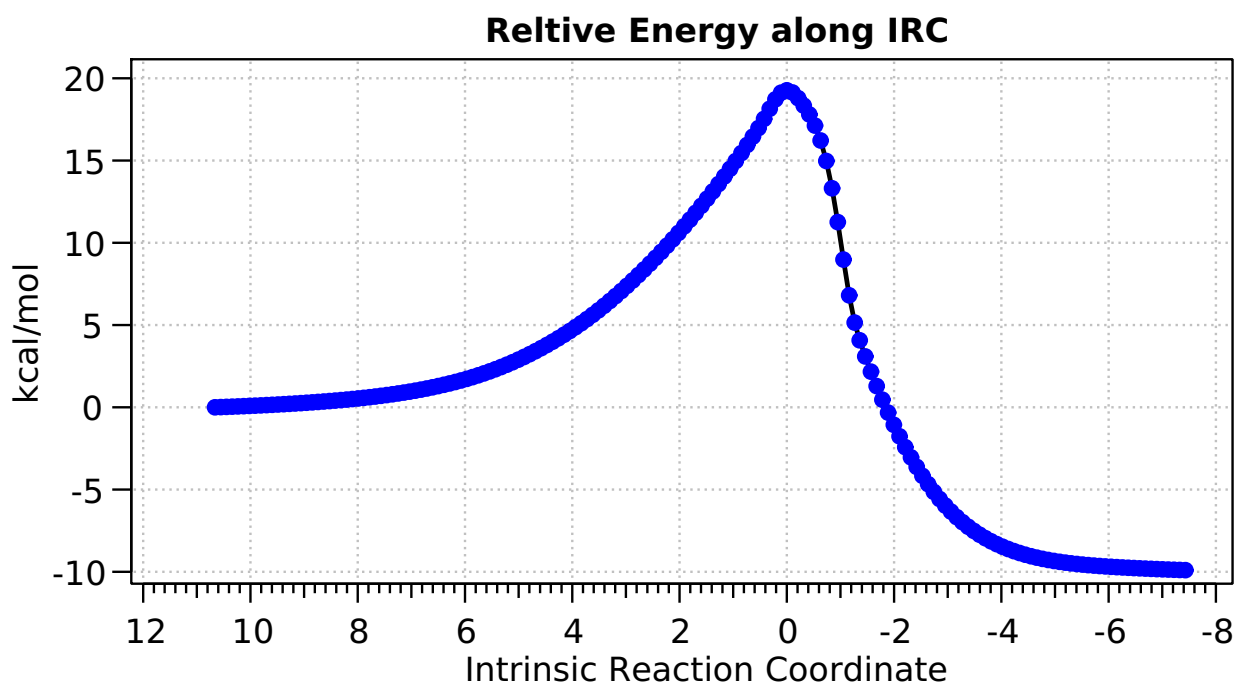
negative peaks represent a species which is being replaced by one giving rise to positive peaks. It is the nature of these latter peaks which directly prove that the species formed after 12 hours contains an asymmetric tetrahedral carbon atom.

1. The negative peaks are  $\sim 61.8$  ppm (triplet), 3.8 (doublet [assigned to the  $\text{CH}_2\text{SH}$  region], 2.6 ppm (the Me group of the acetyl cyanide) and 7.3 (the Ph group).
2. The positive peaks are 1.9 ppm (Me singlet), an AB quartet at 4.2 ppm equal in integration to the negative peak at 3.8 ppm, and a more disperse Ph peak.
3. If pressure is applied [cite]10.1039/P29870001477[/cite] to the solution, the tetrahedral intermediate forms more quickly, and actually crystallises out of solution (Le Chatelier's principle).
4. If a hydroxylamine ( $\text{RNHOH}$ ) is used instead of a thiol, the tetrahedral intermediate forms much more quickly [cite]10.1021/jo00389a050[/cite], and both N: and O: nucleophilic isomers can be detected [cite]10.1021/jo00389a050[/cite] (by  $^{15}\text{N}$ - $^{13}\text{C}$   $^2\text{J}$  couplings in the case of the former).
5. Particularly significant is that the new product manifests its  $\text{CH}_2$  group as an AB quartet. This must mean that the two methylene protons are diastereotopic (have different chemical shifts and hence the coupling between them can be observed). This cannot be achieved by restricted rotation (atropisomerism) and hence can only arise by the  $\text{CH}_2$  group being close to an asymmetric (chiral) carbon atom bearing four different groups.

The other interesting aspect is whether the tetrahedral intermediate is formed in two stages, the first being S-C bond formation, followed by a proton transfer from S to O, or whether it can form directly in a single concerted step. On the assumption that  $\text{CDCl}_3$  solutions contain traces of water, a synchronous mechanism does seem possible. The water helps transfer the hydrogen from S to O, as shown in the intrinsic reaction coordinate [cite]<http://hdl.handle.net/10042/to-11696>[/cite] animation below (wb97XD/6-311G(d,p)/SCRF=chloroform calculation). Notice how the H firstly leaves the S and hops across to the water, and only then does another H from the water hop over to the carbonyl group. The computed free energy barrier for this process is 22.5 kcal/mol, which is pretty much spot on for a slow thermal reaction.



IRC animation for the formation of a tetrahedral intermediate.



The slow motion capture of a tetrahedral intermediate using substituents finely tuned for the purpose (a relatively non nucleophilic thiol coupled with a relatively poor cyano leaving group) enables us to directly prove its identity using NMR. The same technique as it happens was also used to characterise another iconic intermediate, the carbocation, many years after it too was inferred as a key mechanistic participant.

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